



General Certificate of Secondary Education
2025

Centre Number

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Candidate Number

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GCSE Chemistry

Unit 1

Higher Tier

MV24

[GCM12]

MONDAY 19 MAY, MORNING

Time

1 hour 15 minutes, plus your additional time allowance.

Instructions to Candidates

Write your Centre Number and Candidate Number in the spaces provided at the top of this page.

You must answer the questions in the spaces provided.

Do not write on blank pages.

Complete questions in black ink and use a dark HB pencil for drawings and graphs.

Do not write with a gel pen.

Answer all **five** questions.

Information for Candidates

The total mark for this paper is 80.

Figures in brackets printed at the end of each question indicate the marks awarded to each question or part question.

You may use a scientific calculator.

Quality of written communication will be assessed in Question **2(b)**.

A Data Leaflet, which includes a Periodic Table of the Elements, is included in this question paper.

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(Questions continue overleaf)

1 Magnesium reacts with oxygen to form magnesium oxide.

(a) Write a balanced symbol equation for this reaction. [3 marks]

(b) The table opposite shows some properties of magnesium, oxygen and magnesium oxide.

(i) Complete the table opposite. [5 marks]

(ii) Complete the following for a **molecule** of oxygen. [3 marks]

Total number of electrons shown:

Number of lone pairs of electrons:

Number of electrons in covalent bonds:

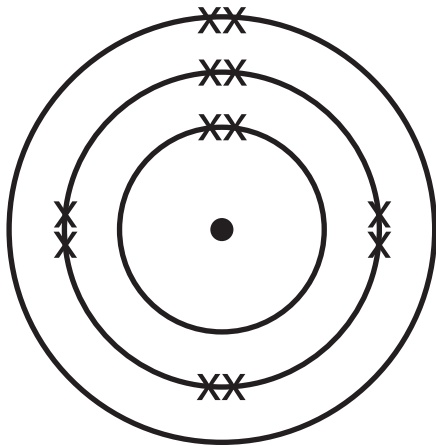
Physical property	Magnesium	Oxygen	Magnesium oxide
Melting point /°C	650	-218	2852
Boiling point /°C	1090	-183	3600
State at room temperature (25°C)	solid	gas	
Colour at room temperature (25°C)	grey	colourless	
Structure			giant lattice
Bonding		covalent	ionic

(iii) Explain why oxygen has a low boiling point. [2 marks]

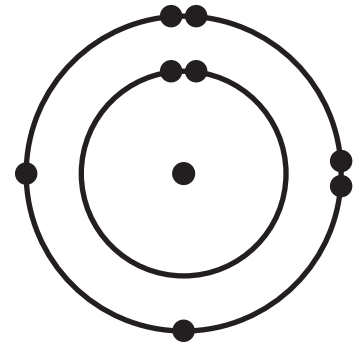
(iv) Explain what is meant by ionic bonding. [1 mark]

(v) Complete the dot and cross diagram below by drawing the electronic configurations of the ions formed and write the formulae of the ions.

[4 marks]



magnesium atom



oxygen atom

magnesium ion

oxide ion

Formula _____

Formula _____

(c) Magnesium oxide is also formed when magnesium reacts with carbon dioxide. The other product is carbon.

(i) Write a balanced symbol equation for this reaction. [3 marks]

(ii) Draw a dot and cross diagram to show the bonding in a molecule of carbon dioxide. [1 mark]

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(Questions continue overleaf)

(a) Suggest one reason why Mendeleev placed zinc in the same group as calcium and magnesium. [1 mark]

(c) Using the Periodic Table in your Data Leaflet, identify the elements below from the information given.

(i) A transition metal which forms a black oxide and a green carbonate.
[1 mark]

(ii) The most reactive halogen.
[1 mark]

(iii) The element in Period 4 and Group 5.
[1 mark]

(iv) The element with atomic number 32.
[1 mark]

(d) Group 1 and Group 7 elements and their compounds undergo a range of chemical reactions.

(i) State the name given to Group 1.
[1 mark]

(ii) Complete the balanced symbol equations opposite for the three reactions shown. Include state symbols. [4 marks]



(iii) When potassium reacts with water in Reaction 1, heat is released and the potassium eventually disappears with a crackle forming a colourless solution.

State three other observations made when potassium reacts with water.

[3 marks]

1. _____
2. _____
3. _____

(iv) State the colour change observed in the solution in Reaction 2. [1 mark]

(v) State what is observed in Reaction 3. [1 mark]

3 Acids react with alkalis and with metal carbonates to form salts.

(a) The reactions of acids occur because of the presence of hydrogen ions in solution.

(i) Write the formula of a hydrogen ion.
[1 mark]

(ii) **Name** the ion present in all alkalis.
[1 mark]

(iii) Sulfuric acid is a strong acid. What is meant by the term **strong** acid?
[1 mark]

(b) The table below shows some acidic and alkaline solutions.

	Solution	Concentration /mol/dm³
A	Ethanoic acid	0.1
B	Hydrochloric acid	0.05
C	Hydrochloric acid	0.1
D	Sodium hydroxide	0.1
E	Sulfuric acid	0.2

(i) Which one of the solutions (**A**, **B**, **C**, **D** or **E**) would have the highest pH? [1 mark]

(ii) Which one of the solutions
(**A**, **B**, **C**, **D** or **E**) would have the
lowest pH? [1 mark]

(iii) What colour is observed when
phenolphthalein is added to
solution **D**? [1 mark]

(iv) Write a balanced symbol equation for
the reaction which occurs between
solution **C** and solution **D**.
[2 marks]

(c) The salt, magnesium sulfate, may be prepared using the steps shown in the flow scheme below.

Step 1 Measure 25cm³ of sulfuric acid and place in a conical flask



Step 2 Add spatula measures of magnesium carbonate until no further reaction occurs



Step 3 Filter the solution



Step 4 Heat the filtrate until the volume of the solution reduces by half



Step 5 Leave aside to cool and crystallise



Step 6 Filter off the crystals and dry them in a low temperature oven

(i) Write a balanced symbol equation for the reaction of magnesium carbonate with sulfuric acid. [2 marks]

(ii) What is meant by the term salt? [2 marks]

(iii) What piece of apparatus is used to measure out 25 cm^3 of sulfuric acid in Step 1? [1 mark]

(iv) How would you know that no further reaction was occurring in Step 2? [1 mark]

(v) Name the three pieces of apparatus required to filter the solution in Step 3.
[1 mark]

1. _____
2. _____
3. _____

(vi) Explain why the solution is heated to half volume in Step 4. [1 mark]

(vii) State two other methods which could be used to dry the crystals in Step 6.
[2 marks]

1. _____
2. _____

4 Aluminium forms a variety of compounds with non-metals.

(a) In an experiment to determine the empirical formula of an aluminium compound, 0.81 g of aluminium reacted with sulfur and formed 2.09 g of a sulfide of aluminium.

(i) Calculate the mass of sulfur which reacted. [1 mark]

mass of sulfur = _____ g

(ii) Calculate the number of moles of sulfur which reacted. [1 mark]

moles of sulfur = _____

(iii) Calculate the number of moles of aluminium which reacted. [1 mark]

moles of aluminium = _____

(iv) Using your answers to **(a)(ii)** and **(iii)**, determine the empirical formula of the sulfide of aluminium. [1 mark]

empirical formula _____

(b) Aluminium oxide reacts with magnesium according to the equation:



459 g of aluminium oxide are mixed with 288 g of magnesium. The mixture is allowed to react.

(i) Complete the table below. [1 mark]

	Aluminium oxide	Magnesium
Mass /g	459	288
Relative formula mass (M_r)	102	24
Number of moles	4.5	

(ii) Which reactant is the limiting reactant? [1 mark]

(iii) Calculate the mass of aluminium formed. [2 marks]

mass of aluminium = _____ g

(c) A sample of 1.5 g of hydrated aluminium nitrate, $\text{Al}(\text{NO}_3)_3 \cdot 9\text{H}_2\text{O}$ was heated to remove all of the water of crystallisation.

(i) What is meant by the term water of crystallisation? [1 mark]

(ii) How could you ensure that all of the water of crystallisation was removed? [1 mark]

(iii) Calculate the loss in mass.
[3 marks]

loss in mass = _____ g

5 (a) The table opposite gives information about the atomic structure of five different atoms and ions.

Complete the table opposite.
[5 marks]

Atom/ Ion	Atomic number	Mass number	Number of protons	Number of neutrons	Number of electrons	Electronic configuration
Li	3	7				
F ⁻		19			10	2, 8
	1	1				2
Ca ²⁺				20		
	7			7		2, 8

(b) Some elements exist as isotopes.

A sample of zinc contains 3 isotopes, details of which are shown in the table below.

Isotope	Abundance /%
^{64}Zn	50
^{66}Zn	25
^{68}Zn	25

Calculate the relative atomic mass (A_r) of zinc. Give your answer to 1 decimal place. [1 mark]

relative atomic mass = _____

**This is the end of the
question paper**

For Examiner's use only	
Question Number	Marks
1	
2	
3	
4	
5	
Total Marks	

Examiner Number

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